

Supporting Information

Turning Carbon Dioxide into Sustainable Food and Chemicals: How Electrosynthesized Acetate Is Paving the Way for Fermentation Innovation

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Supplementary Discussion 1: Electrochemical acetate production model

The techno-economic simulation of electrochemical CO₂ reduction to acetate was performed using our previously published models.^{1–4} To better capture the true cost of electrolysis, a few adjustments were made to the previous input parameters (Table S2). A numerical evaluation of the product price needed for a net present value of 0 at the end of the plant life was used to identify the electrochemical production cost of acetic acid. The modeled performance parameters and materials were based upon that which was experimentally demonstrated in our previous work (Table S3). The downstream separations which included distillation, acetate protonation, and electrolyte recovery were modeled in ASPEN Plus software with the Economic Analyzer plugin as described in our previous work.³ The cost of the CO feed was based upon state-of-the-art high-temperature solid-oxide electrolyzer cell (SOEC) technology. The Topsoe SOEC for CO production was modeled at scale by incorporating published information on the system into our model for CO₂ electrolysis (Table S4). Once the entire model was completed including the high-temperature solid oxide CO₂ electrolysis, low-temperature CO electrolysis, and separation steps, the annualized return on investment (ROI) was determined across a spectrum of production costs and market prices (Equation S1):

$$\text{Annualized ROI} = \frac{\text{Total Profit Generated Over Plant Lifespan}}{\text{Capital Cost Investment} \times \text{Plant Lifespan}} \times 100\% \quad (\text{Eq S1})$$

Supplementary Discussion 2: Techno-economic assessment of electrochemical acetate as a fermentation carbon-source

Based upon the acetic acid production model, it was determined that electrochemical acetic acid can be produced at a cost of 0.32 USD kg⁻¹. The price of fermentation-grade glucose for use in

industrial processes has been previously modeled at 0.40 USD kg⁻¹.⁵ The cost per mol of carbon contained in a molecule of glucose and acetate could then be calculated operating under the assumption that a similar yield could be achieved with both carbon sources during fermentation. Based on this analysis, the cost per mol of carbon was determined to be 20% cheaper from acetate than glucose. The fraction of the total cost made up by the carbon source in industrial scale fermentation processes has been shown to be 77.8%.⁶ Thus, it was determined that the use of acetate as a carbon source instead of glucose can yield a 15.6% lower production cost (Equation S2):

$$\frac{\text{Cost of Carbon Source}}{\text{Total Process Cost}} \times \frac{(\text{Glucose Cost/mol C}) - (\text{Acetate Cost/mol C})}{\text{Glucose Cost/mol C}} \quad (\text{Eq S2})$$

$$\times 100\% = 15.56\%$$

Table S1. Global market size and market price for example target chemicals produced via fermentation.

| Chemical | Market Price (USD kg ⁻¹) | Source | Global Market Size (B USD) | Source |
|-----------------|--------------------------------------|--------|----------------------------|--------|
| Sorbitol | 0.8 | 7 | 1.25 | 8 |
| Erythritol | 3.0 | 9 | 0.20 | 10 |
| Xylitol | 3.0 | 11 | 0.36 | 12 |
| Glycerol | 3.0 | 13 | 2.6 | 14 |
| 2,3-Butanediol | 2.5 | 15 | 0.07 | 16 |
| Citric acid | 4.0 | 17 | 1.75 | 18 |
| Itaconic acid | 1.8 | 19 | 0.12 | 20 |
| Succinic Acid | 2.5 | 15 | 0.13 | 21 |
| Propionic Acid | 1.8 | 22 | 1.6 | 23 |
| Gluconic Acid | 1.4 | 24 | 0.06 | 25 |
| 2-Phenylethanol | 4.3 | 26 | 0.26 | 27 |
| Cobalamin | 30 | 28 | 0.31 | 29 |
| Acetoin | 40 | 15 | 0.25 | 30 |
| Malic Acid | 2.0 | 15 | 0.18 | 31 |
| Fumaric Acid | 1.5 | 15 | 0.67 | 32 |

Table S2. Input parameters for electrolyzer techno-economic assessment.

| Parameter | Value | Source |
|----------------------------|-----------------------------------|--------|
| Plant Production Scale | 50,000 kg day ⁻¹ | 2 |
| Plant Lifetime | 20 years | 2 |
| Income Tax Rate | 38.9% | 2 |
| Nominal Interest Rate | 3.7% | 33 |
| Balance of Plant | 61% of electrolyzer capital cost | 2 |
| Maintenance Cost | 2.5% of electrolyzer capital cost | 34 |
| Material Replacement Costs | 40% of electrolyzer capital cost | 34 |

Table S3. Performance parameters and material usage based upon state-of-the-art CO electrolysis.

| Parameter | Value | Source |
|---------------------|---|--------|
| Cell Voltage | 2.1 V | 3 |
| Faradaic Efficiency | 65% | 3 |
| Current Density | 300 mA cm ⁻² | 3 |
| Cathode Catalyst | 500 mg Cu cm ⁻² | 3 |
| Anode Catalyst | 2 mg Ni cm ⁻² / 1 mg Fe cm ⁻² | 3 |
| Copper Cost* | 1.18 USD kg ⁻¹ | 35 |
| Nickel Cost* | 2.49 USD kg ⁻¹ | 36 |
| Iron Cost* | 0.07 USD kg ⁻¹ | 37 |

*Bulk metal costs were determined by taking a 5-year average from 2015-2019.

Table S4. Topsoe SOEC techno-economic parameters.

| Parameter | Value | Source |
|----------------------|----------------------------|--------|
| Energy Usage | 6 kWh Nm ⁻³ CO | 38 |
| Power Usage | 13 kW m ⁻² | 39 |
| CAPEX | 1,250 USD kW ⁻¹ | 39 |
| Current Density | 850 mA cm ⁻² | 39 |
| Cell Voltage | 1.5 V | 39 |
| CO ₂ Cost | 35 USD ton ⁻¹ | 40 |

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